Common Ionic Charges Reference Page

Ionic compound = Metal + Nonmetal

Element Cations

Element	Charge
Group 1 AND Silver	+1
Group 2 AND Zinc & Cadmium	+2
Group 13 (metals): Aluminum, Gallium, Indium	+3
All other Metals Roman Numerals give the charge (I) +1 (VI) +6 (II) +2 (VII) +7 (III) +3 (VIII) +8 (IV) +4 (IX) +9 (V) +5 (X) +10	VARIES: Roman Numerals show the charge.

Element Anions

Element	Charge
Group 17 Halogens	– 1
Group 16 nonmetals: Oxygen, Sulfur, Selenium	– 2
Group 15 nonmetals: Nitrogen, Phosphorous	– 3

Common Acids

Hydrochloric HCl	Carbonic H_2CO_3
Nitric HNO_3	Sulfuric H_2SO_4
Acetic CH_3COOH	Phosphoric H_3PO_4

Polyatomic Ions

+1	– 1	– 2	– 3
Ammonium NH_4^+	Acetate $C_2H_3O_2^{-1}$	Carbonate CO_3^{-2}	Arsenate AsO_4^{-3}
	Bicarbonate HCO_3^{-1}	Chromate CrO_4^{-2}	Borate BO_3^{-3}
	Bisulfate HSO_4^{-1}	Dichromate $Cr_2O_7^{-2}$	Phosphate PO_4^{-3}
	Bromate BrO_3^{-1}	Peroxide O_2^{-2}	
	Chlorate ClO_3^{-1}	Selenate SeO_4^{-2}	
	Cyanide CN^{-1}	Silicate SiO_3^{-2}	
	Hydroxide OH^{-1}	Sulfate SO_4^{-2}	
	Iodate IO_3^{-1}	Thiosulfate $S_2O_3^{-2}$	
	Nitrate NO_3^{-1}		
	Perchlorate ClO_4^{-1}		
	Permanganate MnO_4^{-1}		
	Thiocyanate SCN^{-1}		

Covalent Compounds

The number of each atom is given by prefixes

Mono-	1
Di-	2
Tri-	3
Tetra-	4
Penta-	5
Hexa-	6
Hepta-	7
Octa-	8
Nona-	9
Deca-	10